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TABLE 17.2 Signs of Enthalpy, Entropy, and Free-Energy Changes and Reaction Spontaneity for a Reaction at Constant Temperature and Pressure

ΔH	ΔS	$\Delta G = \Delta H - T\Delta S$	Reaction Spontaneity
-	+	-	Spontaneous at all temperatures
-	-	- or +	Spontaneous at low temperatures where ΔH outweighs $T\Delta S$ Nonspontaneous at high temperatures where $T\Delta S$ outweighs ΔH
+	-	+	Nonspontaneous at all temperatures
+	+	- or +	Spontaneous at high temperatures where $T\Delta S$ outweighs ΔH Nonspontaneous at low temperatures where ΔH outweighs $T\Delta S$

GIBBS FREE ENERGY

Reactions with a **positive** ΔH and **positive** ΔS all have a negative ΔG

$$\Delta G^\circ = \Delta H^\circ - T\Delta S^\circ$$

$$(+) - T(+)$$

ΔG will be negative in temperature conditions where the value of $T\Delta S$ is higher than the value of ΔH . Thus the reaction is spontaneous only at higher temperatures.



This reaction is **spontaneous** at temperatures above 0°C (273K) and **nonspontaneous** below this temperature

Point	(mol/L)				
1	0.80	334 K	0.00299 K^{-1}	2.048	0.717
2	0.49	321 K	0.00312 K^{-1}	0.471	-0.754
3	0.91	312.1 K	0.00320 K^{-1}	1.43	0.359
4	0.23	303 K	0.00330 K^{-1}	0.0487	-3.023
5	0.11	293 K	0.00341 K^{-1}	0.0879	-4.73

Show sample calculation for K_{sp} and $\ln K_{sp}$ using data point 1 (see equation 6)

Use Excel to plot a graph of $\ln K_{sp}$ (y-axis) vs. $1/T$ (x-axis). Use the "Add Trendline" feature to draw a straight line through the five data points. The equation of the straight line should be shown on the graph along with the value of R^2 . (If one of the data points does not lie on the line, remove it and re-do the straight line, the equation, and R^2 .)

Equation of the straight line $-12889x + 39.362$

Enthalpy of dissolution, ΔH° _____ kJ/mol
Show calculation:

Entropy of dissolution, ΔS° _____ J/mol.K
Show calculation:

Gibbs free energy of dissolution, ΔG° _____ kJ/mol
(at 25°C)
Show calculation:

Name: _____ Lab Day: M T W R F Instructor: _____
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